**Sample Copy of the Lab Report**

1. **Report Cover Page:** *(Fill up the followings, Hand writing)*
2. **Number of the Experiment: ………….**
3. **Name of the Experiment: ………………………………………………………………………………………………………………………………………………………………………………………………………………………..**
4. **Date of Performance: ……………., Date of Submission: …………………………**
5. **Name of Course Teacher: ……………………………………………………………………**
6. **Students’ Name: …………………, ID: …………….., Section: …….., Group: …….**

1. **Body of the Report:** *(Hand writing on A4-size off-set papers)*
2. **PURPOSE/OBJECTIVE**:

………………………………………………………………………………………………………………………………………………………………………………………………………**.***(See Experiment Details/Lab Sheet)*

1. **THEORY:**
2. **Method involved:** (*Acid-base titration/ Redox Titration/ Conductometric Titration)*

……………………………………………………………………………………………………………………………..

**(ii) Reaction:** (Main reactions and Half reactions, if any)

………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………

………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………

**(iii) Indicator:** (Name of the indicator, explain why you have chosen it)

……………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………..

1. **NAME OF THE CHEMICALS:**

**Name of the chemicals Chemical Formula**

*1. …………………………………… ……………………….*

*2. …………………………………… ………………………*

*3. ……………………………………. ……………………….*

*4. ……………………………………. ………………………*

*5. ……………………………………. …………………..etc.*

**[***For example,*

***Name of the chemicals Chemical Formula***

*1. Supplied Sodium Hydroxide solution NaOH*

*2. Standard Oxalic acid solution C2H2O4*

*3. Phenolphthalein indicator C20H14O4*]

1. **NAME OF THE APPARATUS:**

*Burette (50ml) Pipette filler*

*Pipette (10ml) Dropper*

*Conical flask (250ml) Stand clamp etc.*

*Volumetric flask (100ml)*

*Watch glass*

1. **Lab Sheet:** *(Attach the original Lab Sheet signed by your teacher)*
2. **PREPARATION OF APPROX. 0.1N STANDARD SOLUTION:**

The strength of …………. solution =  (N)

1. **EXPERIMENTAL DATA:** *(1 or 2 Tables based on experiment)*

Table-1: ………………………………………………………………………………………….

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| ***No. of reading*** | ***Vol. of ……. ( in ml.)*** | ***Vol. of …….. (burette reading) (in ml.)*** | | | ***Mean (in ml.)*** |
| Initial | Final | Difference |
| 1 | 10 |  |  |  |  |
| 2 | 10 |  |  |  |
| 3 | 10 |  |  |  |
| 4 | 10 |  |  |  |

Table-2: ………………………………………………………………………………………….

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| ***No. of reading*** | ***Vol. of ……. ( in ml.)*** | ***Vol. of …….. (burette reading) (in ml.)*** | | | ***Mean (in ml.)*** |
| Initial | Final | Difference |
| 1 | 10 |  |  |  |  |
| 2 | 10 |  |  |  |
| 3 | 10 |  |  |  |
| 4 | 10 |  |  |  |

1. **CALCULATIONS:**

**………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………..**

1. **RESULTS:**

**………………………………………………………………………………………………………………………..**

1. **Percentage of Errors:** (If necessary)

(Known value – Observed value) × 100

Known value

………………………………………………………..…………………………………………………………………

1. **DISCUSSION:**

**(a) *Precautions Taken*:**

(1) …………………………………………………………………………….

(2) ……………………………………………………………………………

(3) …………………………………………………………………………… etc.

**(b) *Possible errors*:**

(1) …………………………………………………………………………….

(2) ……………………………………………………………………………

(3) …………………………………………………………………………… etc.